

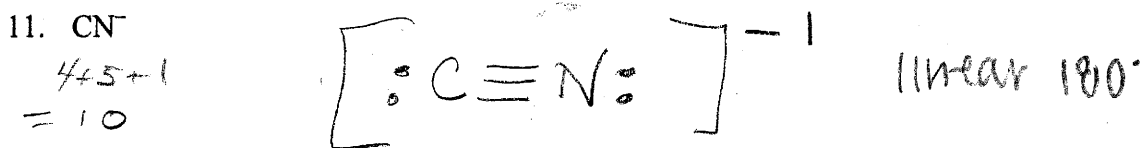
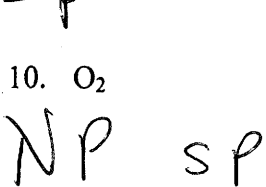
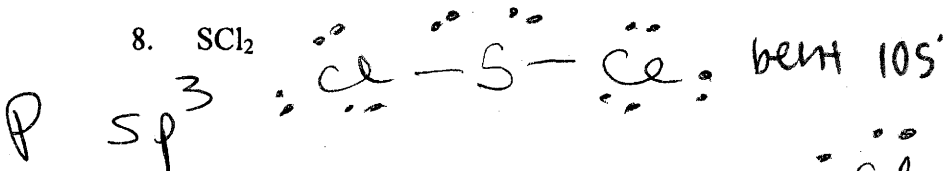
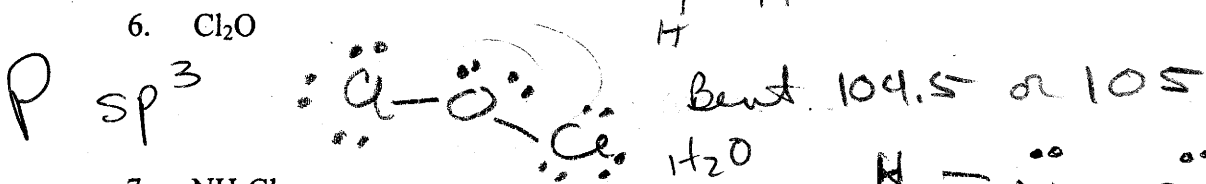
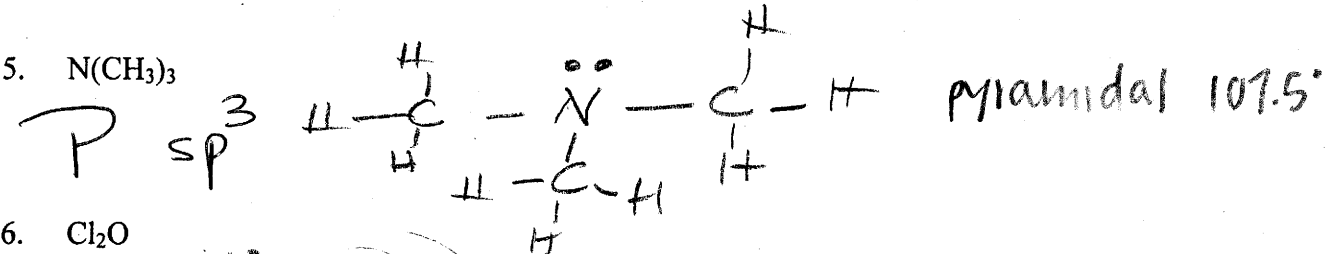
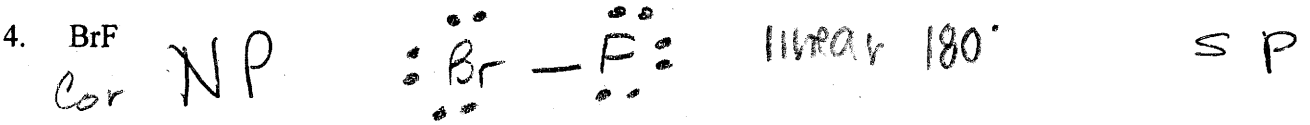
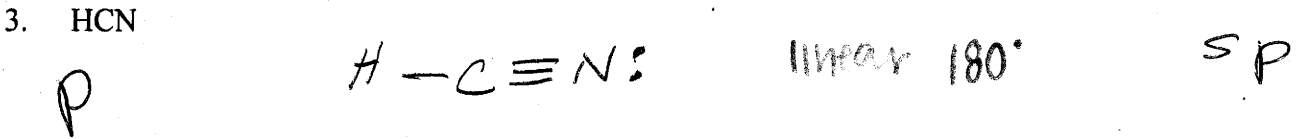
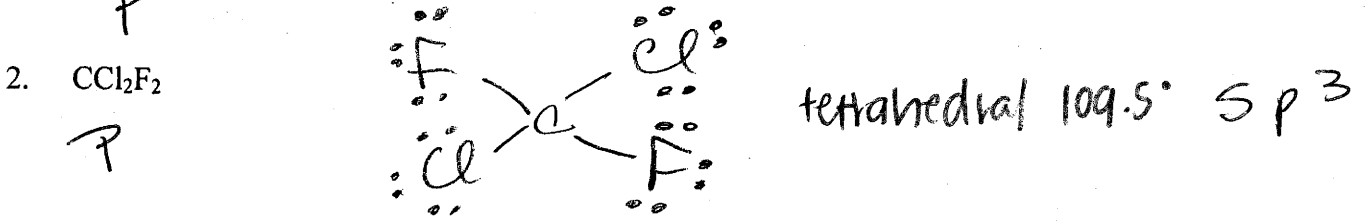
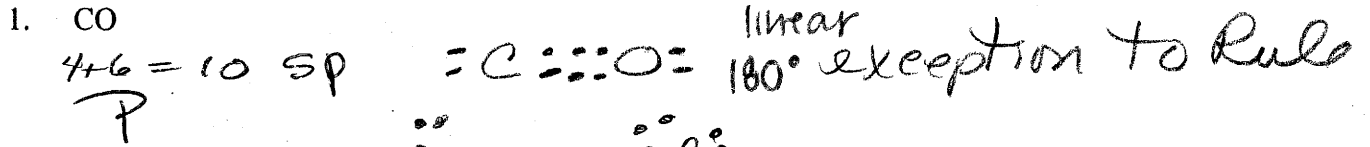
Name: Key

Mixed Practice Lewis Electron Dot Structures

	Name	Ionic, Covalent, Cation, Anion	Chemical Formula	Lewis Structure
1	Copper (II) bromide	Ionic	CuBr ₂	[Cu] ²⁺ 2 [Br:] ⁻
2	Oxide	Anion	O ²⁻	[O:] ²⁻
3	Phosphorus trichloride	Covalent	PCl ₃	Cl:P:Cl:Cl:
4	Dinitrogen tetrafluoride	Covalent	N ₂ F ₄	F ₂ N-F ₂
5	Beryllium (ion)	Cation	Be ²⁺	[Be] ²⁺
6	Dihydrogen monoxide	Covalent	H ₂ O	H-O-H
7	Bromide	Anion	Br ⁻	[Br:] ⁻
8	Francium (ion)	Cation	Fr ⁺	[Fr] ⁺
9	Hydrogen hydroxide	Covalent	HOH	H-O-H
10	Sulfur trioxide	Covalent	SO ₃	O=S=O <i>attached paper</i>
11	Hydrogen monofluoride	Covalent	HF	H-F
12	Dicarbon tetrahydride	Covalent	C ₂ H ₄	H ₂ C=CH ₂
13	Strontium sulfide	Ionic	SrS	[Sr] ²⁺ [S:] ²⁻
14	Sulfide	Anion	S ²⁻	[S:] ²⁻
15	Copper (I) bromide	Ionic	CuBr	[Cu] ⁺ [Br:] ⁻
16	Sulfur dioxide	Covalent	SO ₂	O=S=O
17	Lithium nitride	Ionic	Li ₃ N	3 [Li] ⁺ [N:] ³⁻
18	Magnesium chloride	Ionic	MgCl ₂	[Mg] ²⁺ 2 [Cl:] ⁻
19	Calcium (ion)	Cation	Ca ²⁺	[Ca] ²⁺
20	Chromium (II) fluoride	Ionic	CrF ₂	[Cr] ²⁺ 2 [F:] ⁻

Extra Practice for Electron Dot (Covalent)

Name _____
Date _____ Period _____



P sp all ions Polar