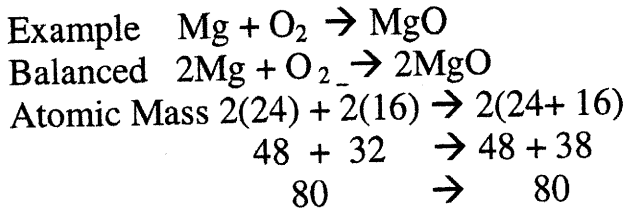
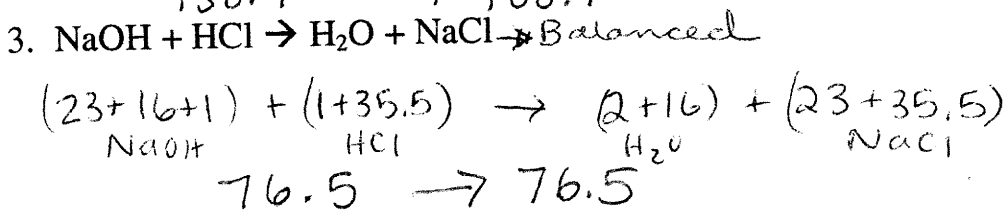
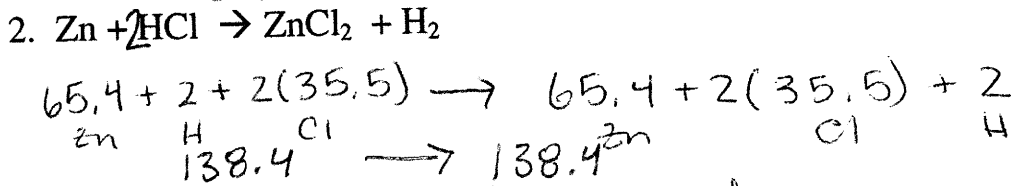
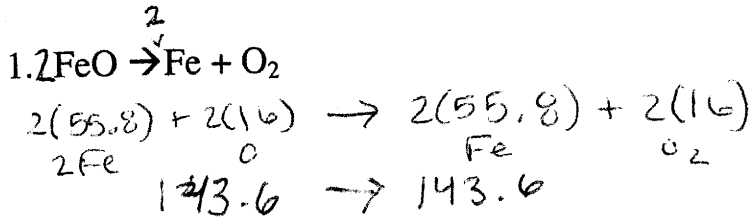


The **Law of Conservation of Mass** states that the mass of all substances before a reaction equals the mass of all substances after a reaction. In other words, the total atomic mass of the reactants should equal the total atomic mass of the products---if the equation is balanced properly. Add the states to each element or compound.

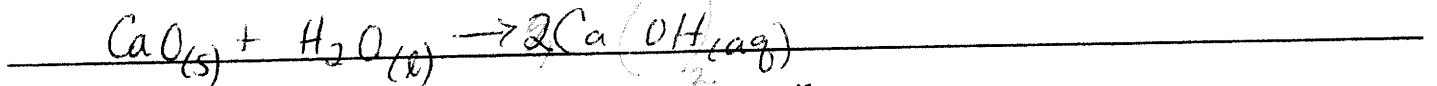


**Part I:** Balance the following equations and using the periodic table, calculate atomic masses of the reactants and products.

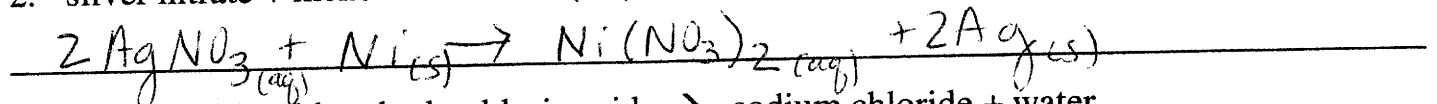


**Part II:** Using the word equation, write and balance the following:

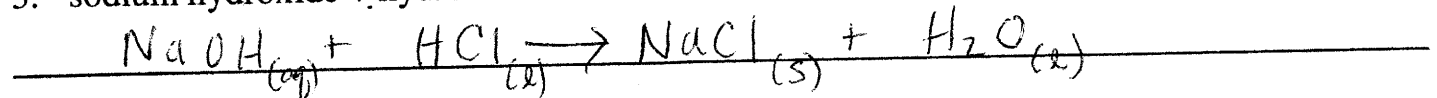
1. calcium oxide + water → calcium hydroxide



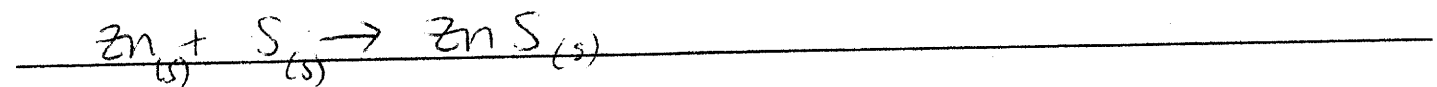
2. silver nitrate + nickel → nickel (II) nitrate + silver



3. sodium hydroxide + hydrochloric acid → sodium chloride + water



4. zinc + sulfur → zinc sulfide



Key 2011  
✓

## BALANCING EQUATIONS PRACTICE

Balance the following equations and include the states (use *s* for metals, *aq* for salts, acids, bases and *g* for gases with parentheses).

