

Oxidation is Loss of e, Reduction is Gain of e

Name _____

Period _____

Redox WS

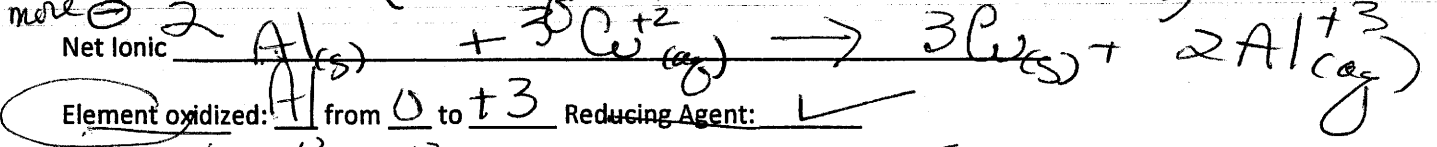
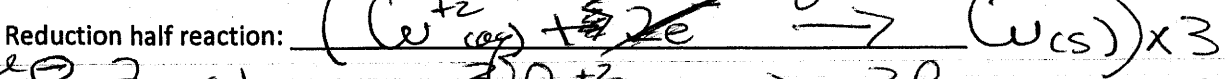
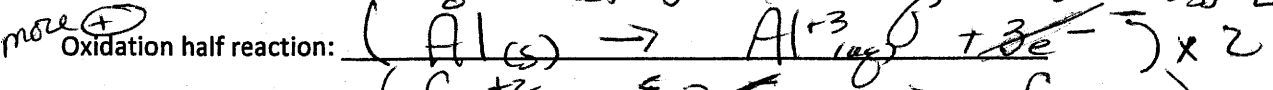
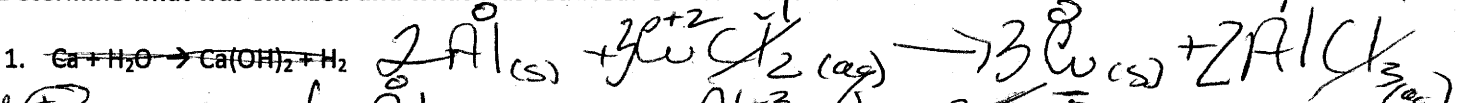
OIL RIG

LEO says GER

Reducing Agent is Oxidized Oxidizing Agent is Reduced

1st Write the oxidation number above each element before and after reaction.

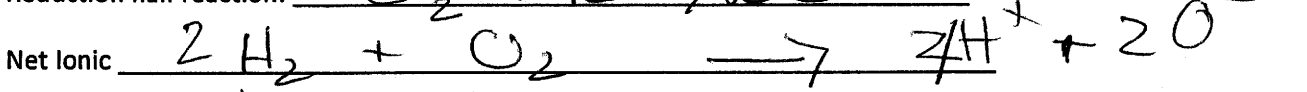
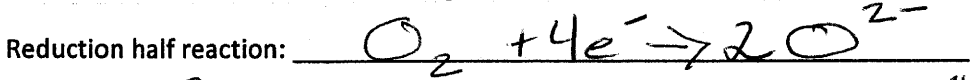
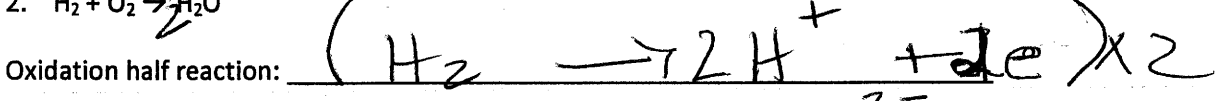
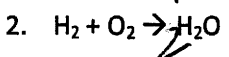
2nd Determine what was oxidized and what was reduced. 3rd Write each half reaction for reactions that occur.



Element oxidized: Al from 0 to +3 Reducing Agent: ✓

Element Reduced: Cu from +2 to 0 Oxidizing Agent: ✓

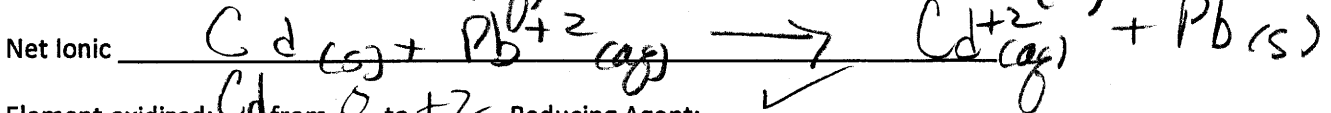
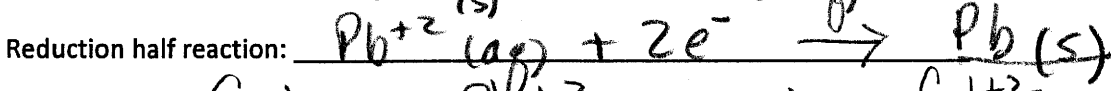
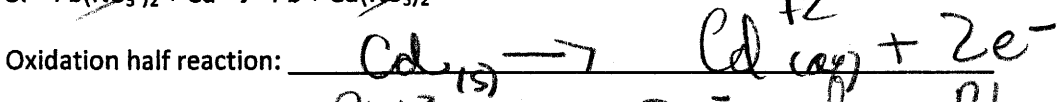
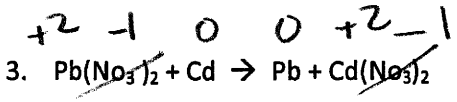
0 0 +1 -2 not on Test



Element oxidized: H from 0 to +1 Reducing Agent: ✓

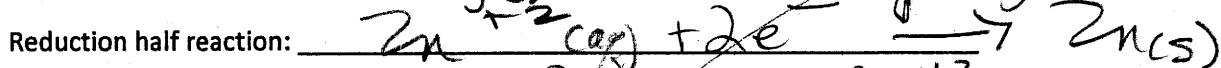
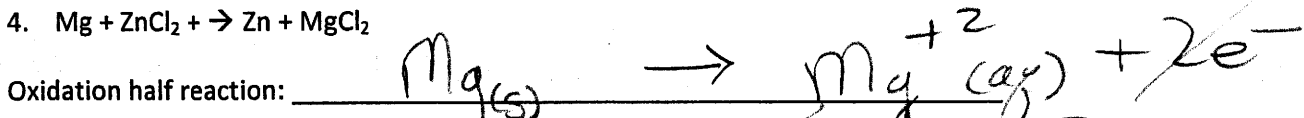
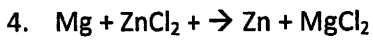
Element Reduced: O from 0 to -2 Oxidizing Agent: ✓

not on Test Quiz



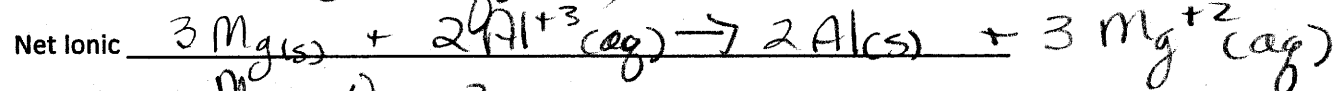
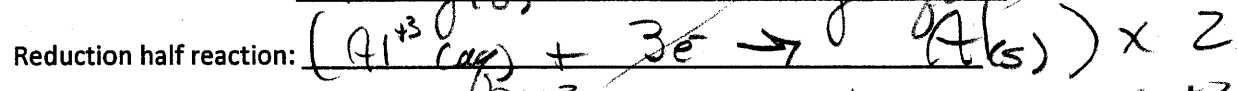
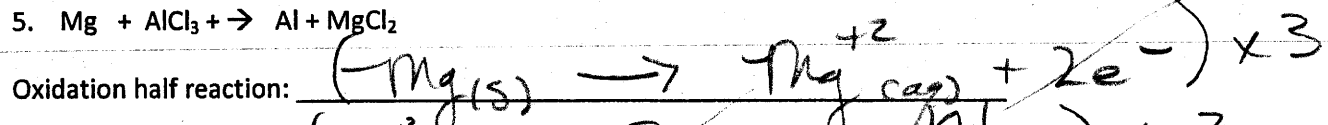
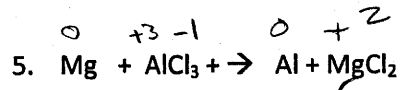
Element oxidized: Cd from 0 to +2 Reducing Agent: ✓

Element Reduced: Pb from +2 to 0 Oxidizing Agent: ✓



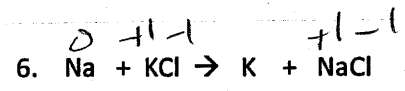
Element oxidized: Mg from 0 to +2 Reducing Agent:

Element Reduced: Zn from +2 to 0 Oxidizing Agent:



Element oxidized: Mg from 0 to +2 Reducing Agent:

Element Reduced: Al from +3 to 0 Oxidizing Agent:



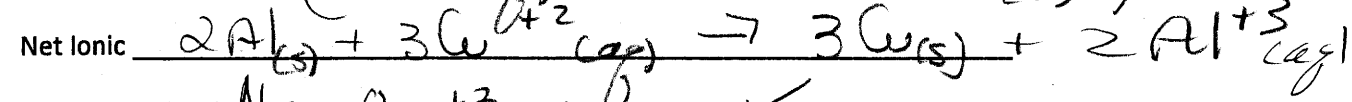
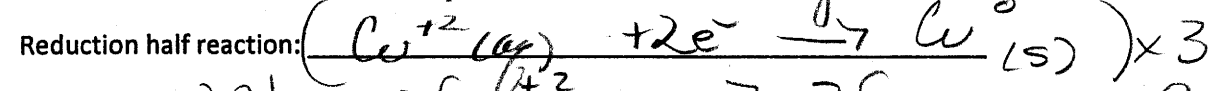
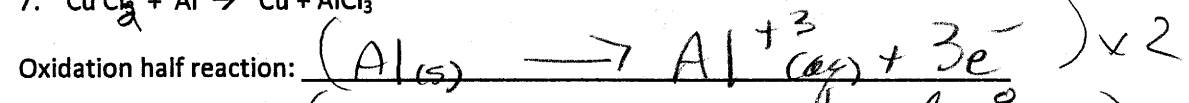
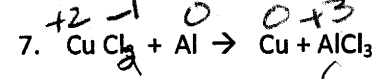
Oxidation half reaction: NR

Reduction half reaction: _____

Net Ionic _____

Element oxidized: ___ from ___ to ___ Reducing Agent: _____

Element Reduced: ___ from ___ to ___ Oxidizing Agent: _____



Element oxidized: Al from 0 to +3 Reducing Agent:

Element Reduced: Cu from +2 to 0 Oxidizing Agent: