

Name KEY 2011

Date _____

Pd. _____

EQUATION PRACTICE

Directions: Write the equation (using formulas), balance the equation, list type of EOC equation.

1. hydrogen gas + nitrogen gas → ammonia



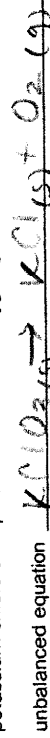
Type of equation: Synthesis EOC #, letter and description 1A: Binary Compd

2. copper + chlorine gas → copper (II) chloride



Type of equation: Synth EOC #, letter and description 1A

3. potassium chlorate → potassium chloride + oxygen gas



Type of equation: Decomp EOC #, letter and description 2E

4. potassium chlorate → potassium chloride + oxygen gas



Balanced equation: _____

Type of equation: _____ EOC #, letter and description _____

5. sodium + water → sodium hydroxide + hydrogen gas



Type of equation: SR EOC #, letter and description 3B

6. potassium chloride + lead (II) nitrate → potassium nitrate + lead (II) chloride



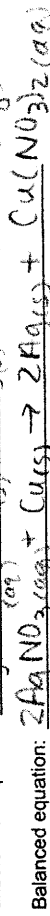
Type of equation: DR EOC #, letter and description 4A

7. sodium hydroxide + iron (III) chloride → sodium chloride + iron (III) hydroxide



Type of equation: DR EOC #, letter and description 4A

8. silver nitrate + copper → silver + copper (II) nitrate



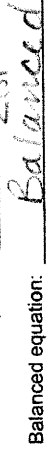
Type of equation: SR EOC #, letter and description 3A

9. magnesium + fluorine gas → magnesium fluoride



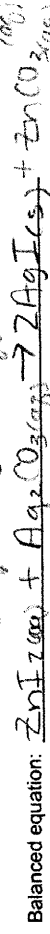
Type of equation: Synth EOC #, letter and description 1A

10. calcium bromide → calcium + bromine gas



Type of equation: Decomp EOC #, letter and description 2A

11. zinc iodide + silver carbonate → silver iodide + zinc carbonate



Type of equation: DR EOC #, letter and description 4A

12. nickel nitrite + copper (I) acetate → copper (I) nitrite + nickel acetate



Type of equation: _____ EOC #, letter and description _____